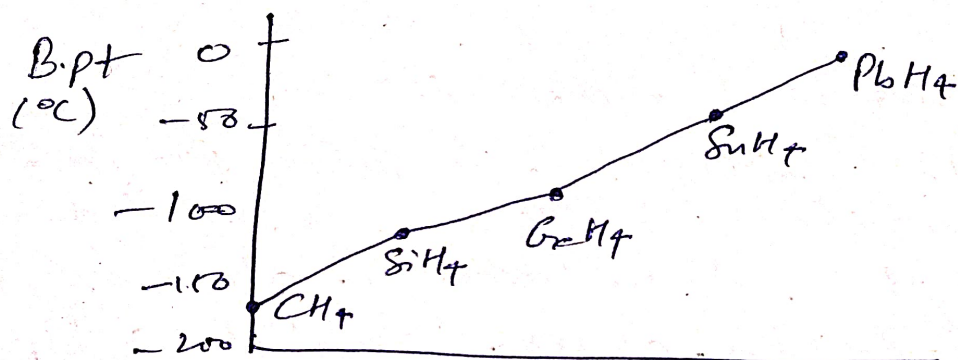


Hydrogen bonding & Weak interactions.

A hydrogen bond is an intermolecular force (IMF) that forms a special type of dipole-dipole attraction when a hydrogen atom bonded to a strongly electronegative atom exists in the vicinity of another electronegative atom with a lone pair of electrons. IMFs occur b/w molecules. Other examples include ordinary dipole-dipole interaction and dispersion forces. Hydrogen bonds are generally stronger than ordinary dipole-dipole & dispersion forces, but weaker than the covalent & ionic bonds.

The evidence for hydrogen bonding



Boiling points of group 14 elemental halides.

The increase in b.pt. happens b/c the molecules are getting larger with more electrons, and so van der Waals dispersion forces become greater.

